

Titration Worksheet

1. What is an acid/base indicator used for?
2. Define "titration":
3. In a few steps, describe how you would titrate a base of unknown concentration with an acid with concentration 1 M.
4. A titration was performed where it took 50.0 mL of 0.10 M hydrochloric acid to neutralize 500.0 mL of a base with unknown concentration. Using this titration information, what was the concentration of the base?
5. A titration was performed where it took 25 mL of 5.0 M NaOH to neutralize 1000.0 mL of an acid with unknown concentration. Using this information, what was the concentration of the acid?